

From Ligand Fields to Molecular Orbitals: Probing the Local Valence Electronic Structure of Ni²⁺ in Aqueous Solution with Resonant Inelastic X-ray Scattering

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Abstract

Bonding of the $\text{Ni}^{2+}(\text{aq})$ complex is investigated with an unprecedented combination of resonant inelastic x-ray scattering (RIXS) measurements and *ab initio* calculations at the Ni L absorption edge. The spectra directly reflect the relative energies of the ligand-field and charge-transfer valence excited states. They give element specific access with atomic resolution to the ground state electronic structure of the complex and allow to quantify ligand field strength and 3d-3d electron correlation interactions in the $\text{Ni}^{2+}(\text{aq})$ complex. The experimentally determined ligand field strength is $10Dq=1.1$ eV. This and the Racah parameters characterizing 3d-3d Coulomb interactions $B=0.13$ eV and $C=0.42$ eV as readily derived from the measured energies match very well with the results from UV/Vis spectroscopy. Our results demonstrate how L-edge RIXS can be used to complement existing spectroscopic tools for the investigation of bonding in 3d transition-metal coordination compounds in solution. The *ab initio* RASPT2 calculation is successfully used to simulate the L-edge RIXS spectra.

Keywords

transition-metal complex, solute-solvent interaction, electron correlation, resonant inelastic x-ray scattering, RASPT2, CFM

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Introduction

3d transition metals in aqueous solution are interesting from a fundamental point of view as their hexaaqua complexes $[\text{M}(\text{H}_2\text{O})_6]^{n+}$ represent the archetype of Werner complexes with coordinative bonding in octahedral symmetry¹. Considering hydrogen bonding between the various coordination spheres of water molecules they could even be regarded as chelate complexes² and new insight into their electronic structure could help understanding the interplay of bonding, structure and dynamics in these complexes. This, in turn, forms the basis for understanding their thermodynamic properties and reactivity, having implications for modeling and understanding phenomena ranging from water exchange reactions^{3,4} to photochemical processes⁵ and atmospheric liquid chemistry of 3d transition-metal ions in solution⁶.

While the structure of 3d transition metal (abbreviated with TM hereafter) ions and their complexes in aqueous solution can be addressed experimentally with numerous scattering and spectroscopic methods⁷, bonding can be assessed with a limited set of spectroscopic tools amongst which, historically, UV/Vis spectroscopy has played a dominant role. Complementing such experimental results and due to the strongly polarized character of the coordinate bond, ligand-field theory^{8,9} can be successfully applied to hexaaqua complexes. In particular, the dependence of the ligand field (LF) state energies on ligand field strength and 3d-3d Coulomb interactions for all $3d^n$ complexes were first calculated by Tanabe and Sugano in their seminal work from 1954^{10,11}. The low energy electronic excited state spectrum of a typical TM coordination complex can thus be described by, with increasing photon energy, the LF excitations, followed by the charge transfer (CT) and ligand-centred excitations. At present, the excitations to LF states of octahedral TM complexes are very well understood at the semi-empirical level of ligand field theory. Still today, ab initio calculations of the LF state energies pose a challenge as exact post-Hartree-Fock or advanced density functional theory (DFT) methods are required to correctly account for static and dynamic correlation effects. Further complications can arise for the calculation of higher energy excited states

such as the CT states. In addition, CT states of TM ions in solution are difficult to probe with UV/Vis spectroscopy compared to LF states. Experimental difficulties arise due to higher transition energies which are well in the UV range and which often strongly overlap with very intense band gap excitations of the host. Thus it appears highly desirable to characterize bonding in aqueous TM ions in general and LF and CT state energies in particular with new experimental tools.

Soft x-ray spectroscopy offers a direct way for probing the valence electronic structure in an element specific way¹². It has the intrinsic property of being sensitive to a local charge distribution of a selected atomic site and, therefore, is ideally suited for studying charge transfer, delocalization and screening processes. Recently, several soft x-ray spectroscopic investigations of the electronic structure of 3d TM ions and compounds in solutions have been reported, based on x-ray absorption spectroscopy (XAS)¹³⁻²², x-ray photoemission spectroscopy (XPS)²³⁻²⁷, resonant photoemission spectroscopy (RPES)²⁴⁻²⁷ and resonant inelastic x-ray scattering (RIXS) also known as resonant x-ray emission spectroscopy (RXES)²⁸⁻³³. Note that RIXS in comparison to XAS and XPS is much less affected by core-hole effects³⁴⁻³⁷. In contrast to XAS and XPS, the resolution of the RIXS spectra is not limited by the core-hole lifetime broadening if sub-natural linewidth excitation is employed^{38,39}. The particular advantage of soft x-ray RIXS for the investigation of TM complexes is that it ideally complements UV/Vis spectroscopy in that it probes the same LF and CT valence-excited states. RIXS is a type of Raman scattering, analogous to the widely used optical Raman scattering. In RIXS the incident photon energy is scanned across the 2p core-level absorption edge of the TM and the energy of the outgoing fluorescence photon is measured. The detected energy losses or energy transfers being the difference of incident and outgoing photon energies directly reflect the energies of the valence excitations. The final states in RIXS are identical to the final states of UV/Vis absorption spectroscopy. In contrast to UV/Vis spectroscopy, RIXS at the TM L-edge is sensitive to local charge distribution in a element specific and symmetry-selective way. In addition, tuning the incident photon energy makes RIXS a truly

two-dimensional spectroscopy, greatly increasing the information content compared to the UV/vis absorption spectroscopy. These features make RIXS a powerful tool for studies of low-energy excitations of complex systems³⁴, therefore making up the added experimental complexity.

In the current paper we investigate electronic structure of the hexaaqua complex $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ with measured and calculated Ni L-edge RIXS. We show how the x-ray spectroscopic results complement information from UV/Vis spectroscopy. We have chosen $\text{Ni}^{2+}(\text{aq})$ for several reasons. First, as the Ni^{2+} -water interaction can be regarded as local and limited to the first solvation sphere only¹⁸ the system can be treated as an octahedral $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ Werner-type coordination complex. Second, with the Ni^{2+} $3d^8$ configuration, multiplet effects arising from core-valence interactions in the Ni^{2+} $2p^53d^9$ intermediate RIXS states are present but expected to be computationally feasible due to the comparably small number of states in contrast to TM ions with less filled 3d shells. The hexaaqua $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex is thus ideally suited to address the interplay of local atomic and intermolecular interactions in hexaaqua complexes. We note that Ni^{2+} compounds have been used before as a test case for XAS and RIXS spectroscopies^{40,41}.

Currently the most successful computational scheme used to simulate the L-edge spectra of TM compounds is the crystal field multiplet (CFM) model developed by B. T. Thole and co-workers^{42,43}. The CFM model is an implementation of the ligand field theory and is thus semi-empirical. Recently a number of theoretical developments have emerged looking for a more general and *ab initio* way to simulate L-edge XAS and RIXS spectra of TM compounds: hybrids of DFT and ligand field theory^{44,45}, 2-hole-2-electron time dependent DFT^{46,47}, DFT-CI^{48,49} and extended Bethe-Salpeter model^{50,51}. Here we present the results of an *ab initio* RASPT2 computational scheme for L-edge RIXS, developed by Josefsson et al.⁵² (a similar method was recently applied in Refs. 32,33,53). Our aim is to evaluate the performance/accuracy of the *ab initio* RASPT2 method with respect to experiment and the well-known CFM model. Note that the CFM model provides an accurate semi-empirical

description of the LF valence states and the core-excited $2p^53d^9$ states. Given the strongly polar Ni^{2+} -water bond, these are expected to dominate the XAS/RIXS spectra. Although for this relatively simple $[Ni(H_2O)_6]^{2+}$ complex the CFM model is sufficient, *ab initio* computational methods have a number of advantages. First, conceptually, an *ab initio* method provides maximum insight to the electronic structure (bonding, electron correlation) and the structure-spectrum relationship. Second, quantum chemistry methods like RASPT2 are implemented for calculations of any molecular system, thus enabling to model systems with low molecular symmetry and/or with complicated or highly covalent bonding interactions, i.e. systems which are difficult to model within semi-empirical ligand field theory framework like CFM.

Methods

Experimental details

$NiCl_2$ salt was purchased from Sigma Aldrich and used without further purification. The salt was solved in deionized water to a concentration of 0.5 mol/l. The dominant species at this concentration is the hexaaqua complex $[Ni(H_2O)_6]^{2+}$ ^{54,55}.

The x-ray absorption (XA) spectrum was measured at the dipole beamline PM3 at the synchrotron radiation source BESSYII of the Helmholtz-Zentrum Berlin in Germany. The spectrum was recorded in transmission mode with a liquid transmission cell set-up¹⁷. The excitation bandwidth was set to 0.4 eV. The photon energy was calibrated using the calibrated water pre-edge peak at 535 eV.

The RIXS spectra were measured at the beamlines U41-PGM and UE52-SGM at BESSYII. The experimental set-up, FlexRIXS, utilizing an in-vacuum liquid jet for sample delivery was used for the measurements³¹. The diameter of the liquid jet amounted to 20 μ m matching the x-ray focus. The x-ray emission spectrometer is mounted at an angle of 90 degrees from the incident x-ray propagation with the polarization of the incident radiation lying in the

scattering plane (linear horizontal polarization) at U41. At UE52-SGM the polarization of the x-ray radiation can be varied freely and this was used to measure spectra both with the polarization axis of the incident radiation horizontal and vertical to the scattering plane. The excitation bandwidth at the Ni L-edge was set to 0.35 eV. The incident photon energy was calibrated with respect to the XA spectrum measured with the liquid transmission cell setup. The x-ray emission spectrometer was used in “slitless” mode using the jet as a source³¹. The resolution of the RIXS spectra was 1.2 eV (determined by the diameter of the jet). The emission energy scale was calibrated using elastically scattered light.

Computational details

The ab initio RASPT2 calculations on $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ were performed using the MOLCAS software package⁵⁶. The geometry of the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex was optimized in the electronic ground state within D_{2h} symmetry at CASPT2 level with an active space consisting of the five nominally Ni 3d orbitals. ANO-RCC basis sets of VTZP quality were used in the CASPT2/RASPT2 calculations^{57,58}. Scalar relativistic effects were introduced via a spin-free Douglas-Kroll-Hess transformation^{59,60}. According to the calculations, the complex is almost perfectly octahedral with a Ni-O distance of 2.080 Å.

Calculations of Ni 2p core-excited states and Ni 3d valence excited states were performed at the RASPT2 level. RASSCF⁶¹ states were calculated with 2p orbitals in the RAS1 subspace (one hole allowed) and 3d orbitals in the RAS2 subspace (all possible permutations allowed). Within each symmetry class in D_{2h} , the states were optimized in a state-averaged RASSCF procedure which was followed by a perturbative inclusion (PT2) of dynamical correlation⁶². Spin-orbit coupling was included using the RASSI technique⁶³. Energies of the core-excited states were shifted by -1.9 eV to compensate for limitations in basis set and active space and match with the experiment. Energies of the valence excitations are presented as calculated.

For comparison we also simulated the RIXS spectra using the semi-empirical CFM model.

The ligand field $10Dq$ parameter was set to 1.1 eV according to the experimentally determined value and the atomic Slater integrals calculated at the Hartree-Fock level were scaled by a factor of 0.8.

For both the ab initio RASPT2 and the CFM calculations, RIXS spectra were simulated by multiplying absorption and emission transition moments without any interference effects included. The following convolution scheme was applied (values for the FWHM are given): 0.35 eV Gaussian broadening taking into account the excitation bandwidth, 0.5 eV (1.0 eV) Lorentzian lifetime broadening at the L_3 (L_2) edge for core-excited states, and 1.2 eV Gaussian broadening of RIXS spectra to account for the spectrometer resolution. An additional 0.5 eV Gaussian broadening of both XA and RIXS spectra was introduced to account for inhomogeneous broadening. No polarization effects were included in the simulation of the RIXS spectra (the calculated RIXS maps thus correspond to excitation with unpolarized light). This is justified as polarization effects on the valence-excited RIXS states were checked to be small (see Supplementary Information) and the Ni^{2+} partial-fluorescence yield XA spectrum derived from the RIXS spectra was shown before to exhibit negligible polarization effects⁶⁴.

Results and Discussion

The main frontier orbitals necessary to characterize the electronic structure of the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex are also the ones probed by Ni L-edge RIXS, namely the nominal water bonding e_g and t_{2g} orbitals and the nominal Ni 3d antibonding e_g^* and t_{2g}^* orbitals (see the Supplementary Information for a molecular orbital diagram and the symmetry labels used here). LF excitations involve only nominal Ni 3d orbitals and correspond to electronic transitions from the ${}^3A_{2g}(e_g, t_{2g})^{10}(t_{2g}^*)^6(e_g^*)^2$ ground state to the final states of the $(e_g, t_{2g})^{10}(t_{2g}^*)^{6-n}(e_g^*)^{2+n}$ ($n=1,2$) configurations. They determine the optical properties of the complex in the UV/visible spectral region. CT excitations involve the nominal water orbitals and correspond to tran-

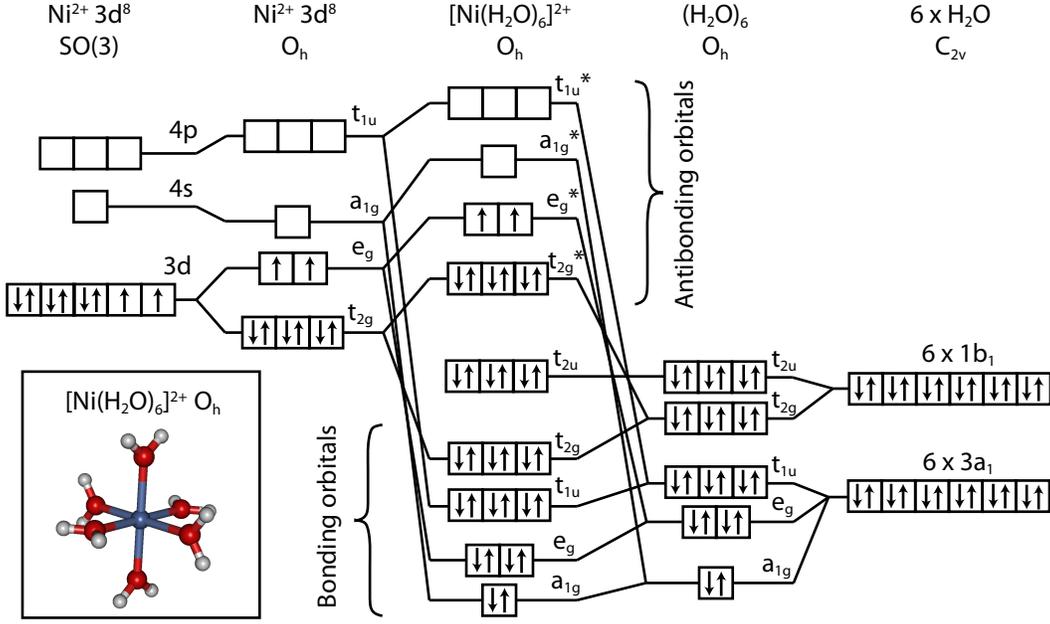


Figure 1: Schematic molecular orbital diagram for the hexaaqua $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex. Antibonding orbitals are marked with *. The calculated structure of the complex is shown in the inset. The orbital labels defined in the figure are used throughout the paper.

sitions from ${}^3A_{2g}(e_g, t_{2g})^{10}(t_{2g}^*, e_g^*)^8$ to states of the $(e_g, t_{2g})^9(t_{2g}^*, e_g^*)^9$ configuration.

With the schematic depiction of the relevant transitions shown in Figure 2(a) we explain how LF and CT states are probed with Ni L-edge RIXS. All energies are plotted with respect to the energy of the ${}^3A_{2g}$ ground state of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ set to zero. Absorption of a soft x-ray photon with incident energy $h\nu_{in}$ promotes the system to a core-excited state with a vacancy in the 2p shell and an additional electron in the unoccupied valence orbitals. The strongest 2p x-ray transitions in $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ occur to the antibonding t_{2g}^* and e_g^* orbitals. The 2p core-excited states in $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ are split by approximately 17 eV due to the 2p spin-orbit interaction separating the L_3 -edge [$(2p_{3/2})^5(t_{2g}^*, e_g^*)^9$ configuration] and the L_2 -edge [$(2p_{1/2})^5(t_{2g}^*, e_g^*)^9$ configuration]. The core-excited states may decay via fluorescence (note that $\sim 99\%$ decay via Auger decay⁶⁵) promoting the system back to the ground- or to valence excited states of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$. By measuring $h\nu_{out}$ the energy transfer $h\nu_{in} - h\nu_{out}$ can be determined which in turn directly reflects the energy of the corresponding LF or CT final states [a RIXS transition to an arbitrary LF final state is shown in Figure 2(a)].

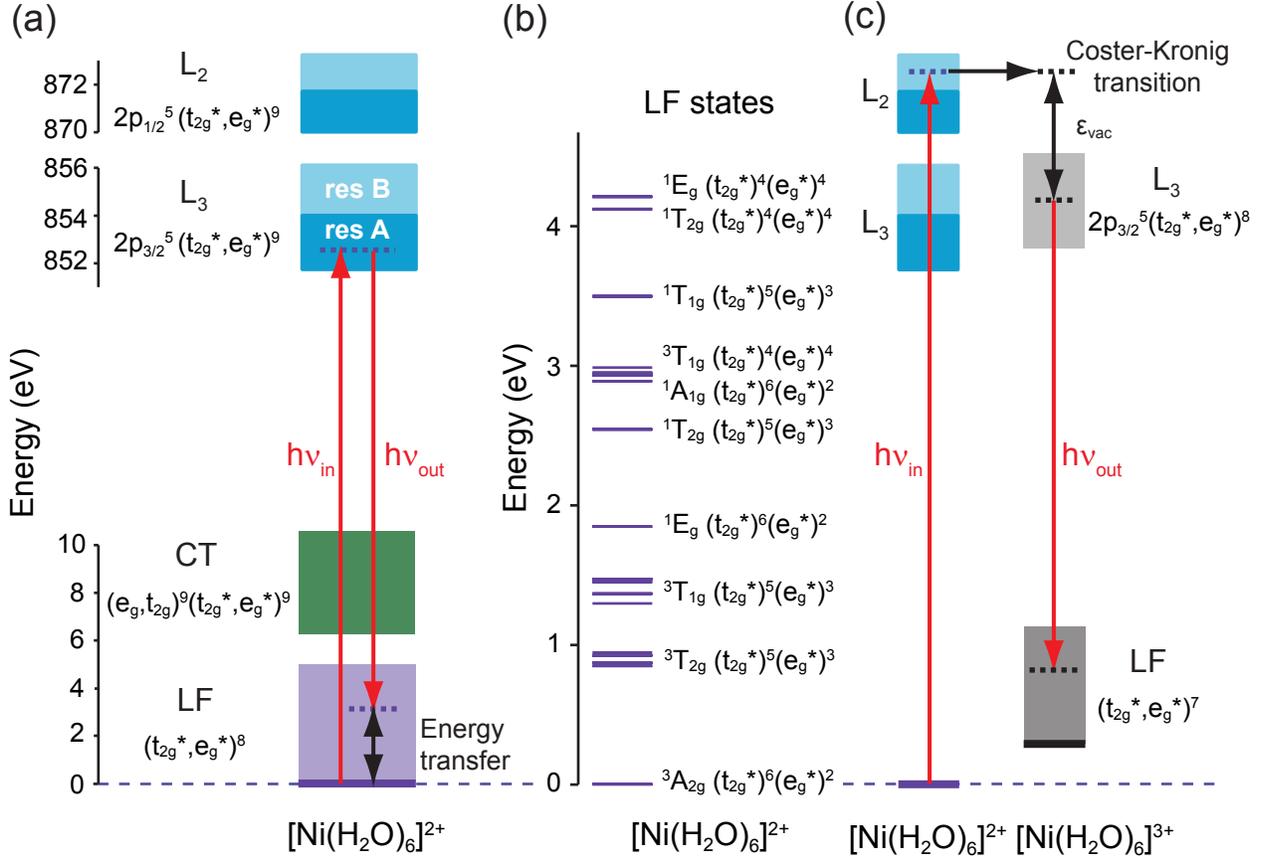


Figure 2: (a) Schematic representation of the relevant transitions and states in L-edge RIXS of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ with approximate relative energies in eV. The filled boxes represent manifolds of closely lying states. The electron configurations of ligand-field (LF) and charge-transfer excited states (CT) and of the core-excited states at the L_3 - and L_2 -edges are given (the two x-ray resonances “res A” and “res B” are defined with Figure 3 and 4). (b) Calculated RASPT2 energy diagram of the LF states of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ denoted by their respective term symbols (the highest-energy LF state $^1A_{1g}$ at 7.3 eV is out of the scale) and the corresponding strong-field orbital configurations (for more details of the orbital and spin compositions see the Supplementary Information). (c) Schematic representation of a selected L_2 -excitation undergoing a Coster-Kronig transition with ejection of an electron (with kinetic energy ϵ_{vac}) and followed by a fluorescence decay with transition to an arbitrary LF state of $[\text{Ni}(\text{H}_2\text{O})_6]^{3+}$. Note that the energy of the ground state of $[\text{Ni}(\text{H}_2\text{O})_6]^{3+}$ is not drawn to scale with respect to the ground state of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$.

The energies of the LF states in $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ as derived from the RASPT2 calculations are plotted in Figure 2(b). Altogether there are 45 states in the LF manifold which can be grouped into 11 terms (states within each term are degenerate in perfect octahedral symmetry and without 3d spin-orbit coupling). The question arises how these LF states can be probed in detail with Ni L-edge RIXS and how this compares to results from UV/Vis absorption spectroscopy.

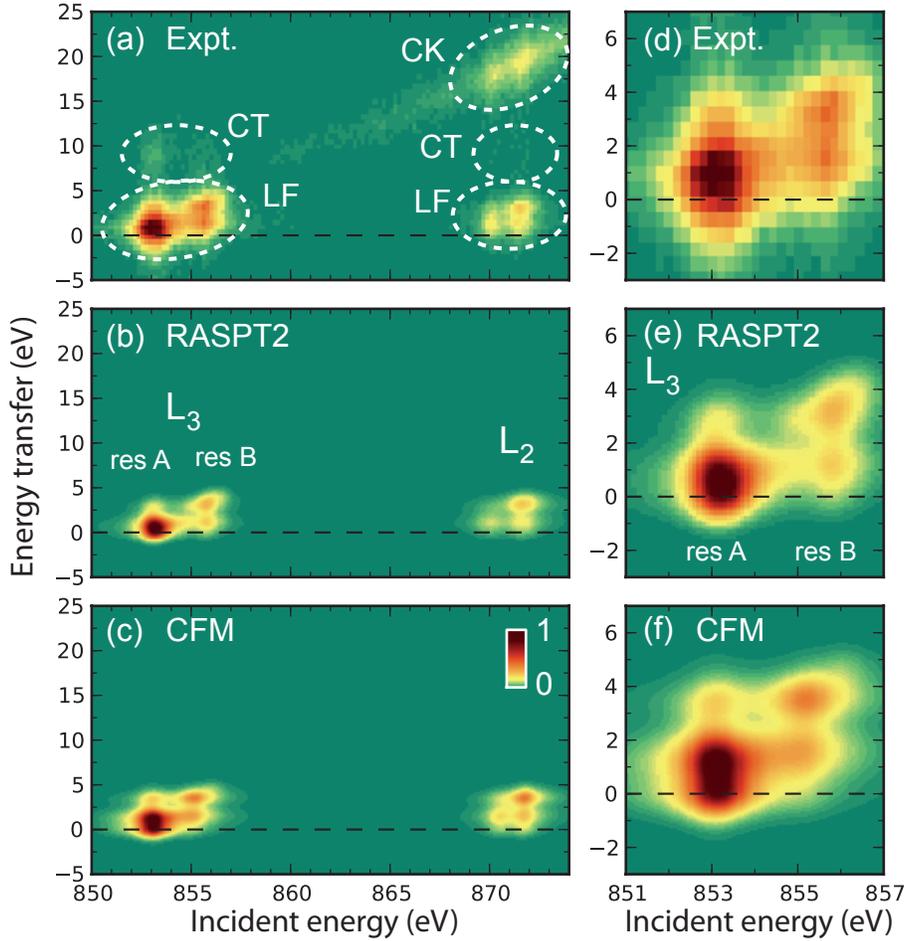


Figure 3: (a)-(c) Measured and calculated Ni L-edge RIXS maps of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ with (a) experiment, (b) RASPT2 calculation, (c) CFM calculation. Spectral features denoted with LF and CT correspond to ligand-field and charge-transfer excited states, respectively. The features denoted CK result from fluorescence decays preceded by Coster-Kronig decays [see schematic diagram in Figure 2(c)]. (d), (e) and (f) are zooms into the LF features at the L_3 edge from experiment, RASPT2 and CFM, respectively. The intensities are normalized to one at maximum and encoded in the color.

For this we inspect the measured and calculated Ni L-edge RIXS results displayed in

Figure 3. The measured data [Figure 3(a) and (d)] are compared to the results of the RASPT2 [Figure 3(b) and (e)] and the CFM [Figure 3(c) and (f)] calculations. We first note the overall very good agreement between experiment and calculations. This is emphasized by the zoomed-in view on the LF states in Figure 3(d), (e) and (f). Inelastic scattering features are apparent at the Ni L_3 - and L_2 -edges at incident photon energies of 851–857 eV and 869–874 eV, respectively. The most intense RIXS features at energy transfers of 0–5 eV at the L_3 -edge correspond to the LF final states. The much weaker intensities at energy transfers of 6–12 eV correspond to ligand-to-metal CT final states. Note that these CT states are not included in the RASPT2 and CFM calculations and accordingly the RIXS features are absent in the calculated RIXS maps in Figure 3(b) and (c). The CT final states are excluded here due to practical reasons – it requires more effort to accurately model both the LF and CT states than only the LF states. Note that the CT states nevertheless carry significant information about the electronic structure, specially related to the covalent orbital mixing. An attempt to model the CT states is made in an upcoming publication⁶⁶. We apply there both CFM and ab initio RASPT2 methods to calculate the CT RIXS feature. For now we will focus on the LF states which is sufficient for discussion of ligand-field and 3d-3d Coulomb interactions (static electron correlation).

The last set of final states can be found at 12–22 eV and at the L_2 -edge only. These intensities are due to a cascade process including a radiationless Coster-Kronig (CK) transition⁶⁷ as schematically depicted in Figure 2(c) and as discussed before for TM ions in solution in Refs. 19 and 30. Here the $2p_{1/2}$ core-excited states decay via a CK process upon electron ejection to $2p_{3/2}$ core-excited states of $[\text{Ni}(\text{H}_2\text{O})_6]^{3+}$ complex. These then decay radiatively to the valence-excited states of the $[\text{Ni}(\text{H}_2\text{O})_6]^{3+}$ complex. In Figure 3 the corresponding features are labeled with CK. As shown in Figure 2(c), the energy transfers of the CK features can be estimated from the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ LF-state energies plus the Ni 2p spin-orbit splitting of 17 eV. Finally we point out that the intensity of elastic scattering at an energy transfer of 0 eV is slightly overestimated in both calculations [compare Figure 3(d), (e) and

(f)]. The intensity of elastic scattering depends on a number of factors such as the experimental geometry, the polarization of the incident radiation and the sample surface shape and roughness. In the simulated RIXS spectra these effects are not included.

The agreement of measured and calculated spectra enables the determination of the ligand field strength and 3d-3d Coulomb repulsion parameters as will be discussed in the following Section, together with a detailed comparison of the two theoretical approaches.

Table 1: Measured and calculated energies of the given LF states with respect to the ${}^3A_{2g}$ ground state. The measured energies result from peak fits to the experimental spectra shown in Figure 5 (see Supplementary Information, Figure S3 and Tabel S1). Due to insufficient resolution some of the LF states could not be resolved in the experiment. The calculated RASPT2 energies are averages for the given term. The calculates RAS energies are taken from Ref. 68.

	RIXS Expt. This work (eV)	RASPT2 Theory This work (eV)	CFM Theory This work (eV)	RAS Theory Ref. 68 (eV)
${}^3A_{2g}$	0	0	0	0
${}^3T_{2g}$	1.1	0.9	1.1	1.0
${}^3T_{1g}(F)$	(1.4)	1.4	1.8	1.7
${}^1E_g(D)$	-	1.8	1.9	2.2
${}^1T_{2g}(D)$	-	2.5	2.9	3.1
${}^1A_{1g}(D)$	-	2.9	3.2	3.5
${}^3T_{1g}(P)$	3.4	2.9	3.4	3.6
${}^1T_{1g}$	3.5	3.5	3.6	-
${}^1T_{2g}(G)$	-	4.1	4.4	-
${}^1E_g(G)$	-	4.2	4.5	-
${}^1A_{1g}(S)$	-	7.3	7.8	-

Determination of ligand field and Racah parameters

As the same final states are reached with excitation at the L_3 - and L_2 -edge albeit with smaller intensities at L_2 , we focus in the following on excitations at the L_3 -edge only. We denote the two clearly distinct resonances at the L_3 -edge as resonance A at 853.2 eV and resonance B at 855.6 eV. States in resonance A have dominantly triplet character (anti-parallel spins of 2p and 3d hole), whereas states in resonance B have considerable singlet character (see

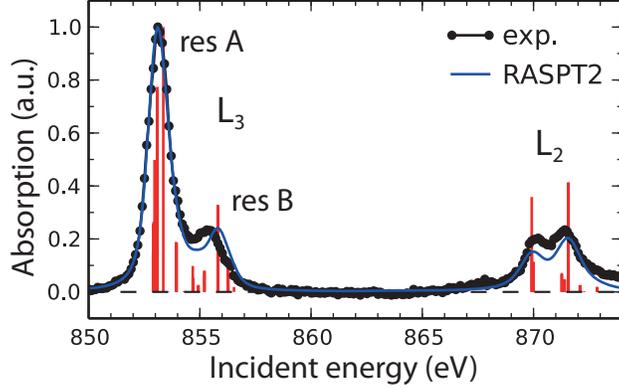


Figure 4: Measured (black circles) and calculated (RASPT2, solid blue line and red sticks representing single transitions) Ni L-edge x-ray absorption spectrum of Ni²⁺(aq). The experimental x-ray absorption spectrum was measured in transmission mode and published before in Ref. 52.

also Figure S2). RIXS spectra displaying intensities versus energy transfer at these two resonances (corresponding to vertical cuts through the maps in Figure 3) are presented in Figure 5. The peak at zero energy transfer in Figure 5(a) corresponds to transitions back to the ground state ${}^3A_{2g}$ (resonant elastic scattering). The intensity of the elastic scattering peak depends on the position of the detector with respect to the polarization plane of the incident radiation. It is enhanced if photons are detected normal to the polarization plane (vertical polarization of the incident radiation in our case, see Methods) as demonstrated in Figure 5(a). The RIXS peak at a transfer of 1.1 eV corresponds to the ${}^3T_{2g}$ final state. The deduced peak energy separation of 1.1 eV with respect to the ground state equals the LF parameter $10Dq^9$. A detailed comparison of calculated RASPT2 and measured RIXS spectra at resonances A and B is done with Figure 5(b) and (c). This comparison demonstrates that by varying the incident photon energy the relative contributions of the various LF states can be varied in the RIXS spectra. At resonance B this enables us to extract the energy of ${}^1T_{1g}$ state. The comparison of experimental and calculated values for the LF states are summarized in Table 1. The experimental LF state energies were retrieved by peak fitting the RIXS spectra in Figure 5 (see also Supplementary Information). The ${}^3T_{2g}$ and ${}^3T_{1g}(P)$ state energies result accordingly from a fit to the RIXS spectrum measured at resonance A

[Figure 5(b) and Figure S3(c)]. The ${}^1T_{1g}$ state energy results from a fit to the RIXS spectrum measured at resonance B [Figure 5(c) and Figure S3(d)]. Note that the lower energy peak in the latter spectrum is a mixture of ${}^3T_{2g}$, ${}^3T_{1g}(F)$ and ${}^1E_g(D)$ [see Figure 5(c)] and can therefore only tentatively be assigned to ${}^3T_{1g}(F)$ (hence the parenthesis in Table 1).

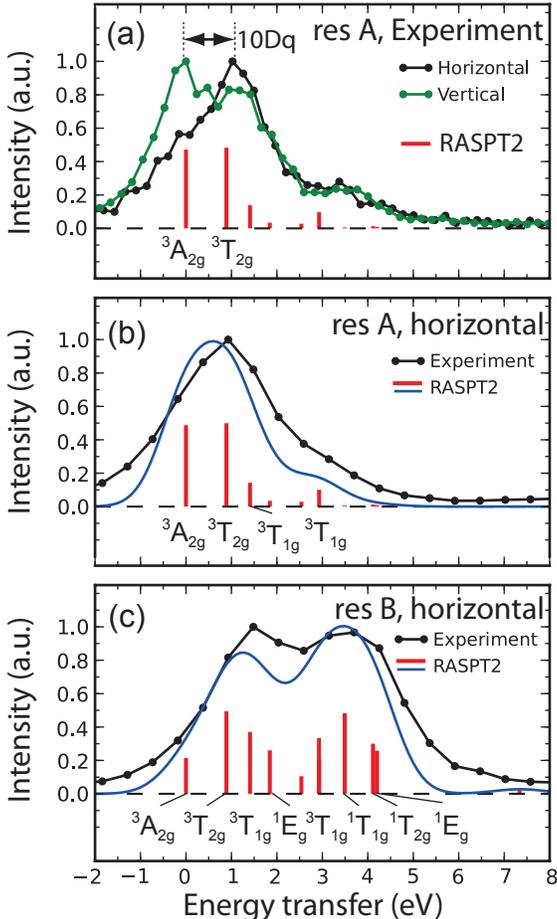


Figure 5: Measured and calculated Ni L_3 -RIXS spectra of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ in the energy-transfer region of the ligand-field (LF) features with: (a) Experimental spectra with excitation at resonance A (incident photon energy 853.2 eV) for horizontal (black) and vertical (green) polarization of the incident radiation (see Methods for details) and calculated LF peaks (RASPT2, red sticks, horizontal polarization). (b) and (c) Experimental spectra (black circles and lines) with excitation at resonance A (853.2 eV) and B (855.6 eV), respectively, compared to the calculated spectra (RASPT2, blue solid lines). The sticks represent individual peaks with corresponding assignments of the LF states. The intensities are normalized to one at maximum.

We note that in spite of the overall good agreement of RASPT2 and CFM spectra with experiment, the RASPT2 calculation tends to underestimate the LF state energies compared

to experiment and compared to the CFM and RAS calculations from Ref. 68 by 10–20%. First, this mismatch of the calculated RASPT2 LF state energies could be attributed to the exclusion of vibrational effects in the calculation. Optical transitions are generally accompanied by vibrational excitations, causing the central energy of experimental peaks to shift towards higher energies. In the UV/vis spectrum of the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex, the peaks are shifted by 0.1 to 0.2 eV compared to pure electronic transitions⁶⁹. One could expect vibrational shifts of similar magnitude in the RIXS spectra explain the discrepancy between RIXS calculation and experiment. There are important differences between the vibrational effects in UV/vis and RIXS spectroscopy, though, which are discussed in the next section. Note that the empirical parameters used in the CFM calculation account for the vibrational effects as calculated peak positions are effectively adjusted to the experimental ones. The second cause for the mismatch between measured and calculated RASPT2 LF state energies could be attributed to the insufficient inclusion of the dynamical correlations by the PT2 method with the current choice of active space. This could also explain the differences between the RASPT2 and the RAS calculations⁶⁸.

Our assignments in Table 1 differ from the results in Ref. 28 which can be explained by the inconsistencies introduced when using one-electron models such as used in Ref. 28 in contrast to the many-electron approach applied here. It is well-known that a many-electron description is necessary in order to interpret the metal L-edge XA and RIXS spectral features of TR compounds^{35,70,71}.

In ligand field theory the LF state energies are determined by $10Dq$ and the Racah parameters B and C (not considering the 3d spin-orbit interaction). In addition to $10Dq$ we can also uniquely determine the values of the Racah parameters. We utilized the following formula from Ref. 9 to calculate B and C (state energies are with respect to the ground

state ${}^3A_{2g}$):

$$B = \frac{E({}^3T_{1g}(P))^2 - 30DqE({}^3T_{1g}(P)) + 200Dq^2}{15E({}^3T_{1g}(P)) - 270Dq}, \quad (1)$$

$$C = 0.5E({}^1T_{1g}(G)) - 6B - 5Dq. \quad (2)$$

For quantitative comparison we also derived these parameters from the RASPT2 and CFM calculations. In the CFM calculation, $10Dq$ and the Coulomb repulsion parameters are actually input parameters and we used here the experimental value of 1.1 eV for $10Dq$ and the Slater integrals F_2 and F_4 were calculated at the Hartree-Fock level and scaled by an empirical factor of 0.8 (the Slater integrals are related to the Racah parameters by the following equations: $F_2 = 49B + C$ and $F_4 = 63/5C$).

As accurate ab initio calculations of the Coulomb repulsion parameters are generally still a challenge we are particularly interested in comparing the values of B and C as derived from the RASPT2 calculation to the experimental and CFM values. The results are summarized in Table 2.

The parameters derived from the RIXS experiment match very well with the UV/Vis results. The ab initio RASPT2 values are off by 10-30%, similarly to the respective energies of the LF states in Table 1.

Table 2: Electronic structure parameters of the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex which determine the energetic positions of the LF states, namely the LF strength parameter $10Dq$ and the Racah parameters B and C , as determined from the present Ni L-edge RIXS experiment and as calculated here with the RASPT2 and CFM approaches. The experimental UV/Vis results are taken from Ref. 69.

	RIXS Expt. This work (eV)	UV/Vis Expt. Ref. 69 (eV)	RASPT2 Theory This work (eV)	CFM Theory This work (eV)	(CFM-RASPT2)/CFM
$10Dq$	1.1	1.1	0.9	1.1	18%
B	0.13	0.12	0.11	0.13	12%
C	0.42	0.47	0.61	0.48	-27%

Comparison of RIXS and UV/Vis absorption spectroscopy

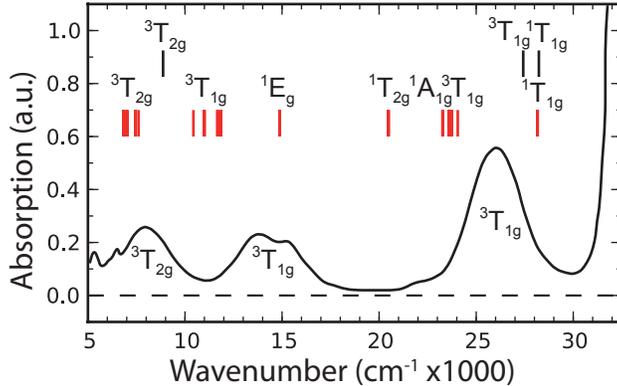


Figure 6: Experimental UV/Vis absorption spectrum from Ref. 69 (black solid line) compared with the energies of the LF states deduced from the L_3 -RIXS experiment (black sticks) and from the RASPT2 calculation [red sticks, same energies and states as in Figure 2(b)].

While in UV/Vis absorption spectroscopy only single dipole transitions contribute to the spectrum, the RIXS process is described by a scattering process (two dipole transitions) through an intermediate state involving the highly localized atomic core orbital. In the optical absorption spectrum transitions to LF states are by a factor of 1/1000 weaker compared to strong CT and ligand-centered absorption peaks⁸. This is because in octahedral TM complexes direct optical dipole transitions between the different LF states are Laporte forbidden and spin-allowed transitions have noticeable oscillator strengths only due to non-octahedral dynamical distortions (vibronic coupling)⁶⁹. In contrast, RIXS transitions to the LF states are allowed in octahedral TM complexes and the RIXS spectrum is typically dominated by the LF features.

The UV/Vis absorption spectrum of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ as adapted from Ref. 69 is displayed in Figure 6. Comparison with the RIXS spectra in Figure 5 reveals some important differences between these techniques. First, starting from the A_{2g} ground state of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ the allowed direct (single dipole) transitions are only to T_{2u} states, whereas the RIXS scattering selection rules allow transitions to final states with A_{2g} , E_g , T_{1g} and T_{2g} symmetry. Second, the spin selection rule is well preserved in UV/Vis absorption. As the 3d spin-orbit coupling

is weak only dipole transitions which preserve the spin have intensity in the UV/Vis spectrum of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ and starting from ${}^3\text{A}_{2g}$ only triplet states are reached (Figure 6). The 2p spin-orbit coupling, however, is the dominant interaction in the core-excited intermediate state in Ni L-edge RIXS (stronger than LF and 2p-3d direct and exchange Coulomb interactions). Therefore, spin is not a good quantum number for the intermediate states and depending on the selected intermediate state, LF final states with different spin multiplicity than the ground state can have similar RIXS intensities. For Ni L-edge RIXS of $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ this allows probing singlet final states. One example for this is the peak at 3.5 eV energy transfer at XA resonance B corresponding to the ${}^1\text{T}_{1g}$ final state [Figure 5(c)]. Third, the possibility to tune the incident photon energy to resonantly select particular intermediate states adds an extra dimension to RIXS as relative intensities of the final states depend strongly on the chosen x-ray absorption resonance. We utilized this here to determine the energies of the closely lying ${}^3\text{T}_{1g}$ and ${}^1\text{T}_{1g}$ LF states at energy transfers of 3.4 and 3.5 eV, respectively. Their separation of only 0.1 eV is well below our experimental resolution. However, tuning to resonances A and B enhances the intensities of ${}^3\text{T}_{1g}$ and ${}^1\text{T}_{1g}$, respectively, as ${}^1\text{T}_{1g}$ cannot be reached when exciting at resonance A and allows for their separate investigation.

Finally, the direct comparison of UV/Vis and RIXS results in Figure 6 shows that the LF-state energies do not exactly match: the energies of the ${}^3\text{T}_{2g}$ and ${}^3\text{T}_{1g}$ states as measured with RIXS are shifted to higher energies with respect to the corresponding UV/Vis absorption peaks. Such an observation was made also in the comparison of UV/Vis and L-edge RIXS results of cobalt compounds in Ref. 72. We believe that these discrepancies can be explained by considering the vibrational effects. The vibrational effects in the UV/Vis spectrum of the $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex were simulated in Ref. 69 and similar to the mismatch of RIXS with the calculations we believe that the blue shift of the RIXS energies compared to the UV/vis energies can be attributed to differences of vibrational effects in the respective methods. Qualitatively one can outline these differences of vibrational effects on spectral peak positions and shapes between UV/Vis and RIXS. First, in case of a nominally octahedral complex,

UV/Vis spectroscopy probes preferably distorted configurations, whereas RIXS in contrast can probe also strictly non-distorted (octahedral) configurations. Secondly, intermediate core-excited states in RIXS are generally vibrationally highly excited and this allows reaching higher vibrational states in the electronic final state compared to UV/Vis spectroscopy. Therefore, extended vibrational progressions appear in the RIXS spectrum and we expect this to lead to larger effective vibrational energy shifts. A quantitative evaluation of these effects would be desirable. However, this requires the calculation of numerous multidimensional potential energy surfaces, including the core-excited states, and the subsequent calculation of nuclear wavefunctions³⁹. The RASPT2 method provides a suitable computational framework for this but it goes beyond the scope of the present investigation.

Conclusions

We investigated electronic structure of the archetypical hexaaqua complex $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ with an unprecedented combination of resonant inelastic soft x-ray scattering (RIXS) measurements and *ab initio* calculations at the Ni L absorption edge. The RIXS spectra directly reflect the relative energies of the ligand-field and charge-transfer valence excited states of the complex and give element-specific access to the ground-state electronic structure of the complex. The results allow for determining the ligand field strength $10Dq$ and 3d-3d Coulomb repulsion in $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$. The $10Dq=1.1$ eV estimate and the Racah parameters $B=0.13$ eV and $C=0.42$ eV characterizing 3d-3d Coulomb interactions as derived from the measured ligand-field-state energies compare very well with the results from UV/Vis spectroscopy. These results show in general how L-edge RIXS can be used to complement existing spectroscopic tools for the investigation of bonding in 3d transition-metal coordination compounds in solution.

We have also demonstrated for $[\text{Ni}(\text{H}_2\text{O})_6]^{2+}$ complex that a full simulation of Ni L-edge RIXS is feasible within a novel RASPT2 computational scheme. Analysis of the RASPT2

results showed that the electronic structure parameters $10Dq$, B and C agree with experiment to within 20%. These discrepancies are mostly a result of the vibrational effects not included in the calculation and not due to errors in calculating the electronic structure. Latter errors might be present only due to insufficient description of the dynamical correlation. The novel RASPT2 calculations are compared in detail to the results of more established semi-empirical CFM calculations. This demonstrates that *ab initio* RASPT2 calculations can be considered as a valuable new tool for the analysis of 3d transition-metal L-edge spectroscopy. Note that the RASPT2 method, based on a LCAO-MO approach, can be used to provide considerably more insight to the bonding of TR-complexes than semi-empirical CFM calculations.

We believe that L-edge RIXS combined with *ab initio* RASPT2 calculations can be utilized to resolve complex questions regarding bonding and chemistry of TR-compounds, therefore contributing to the understanding of their catalytic and photoactive properties.

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Supporting Information Available

Addition details regarding assignments of the experimental and calculated spectral features and determination of the experimental peak positions. This material is available free of charge via the Internet at <http://pubs.acs.org/>.

References

- (1) Werner, A. Beitrag zur Konstitution anorganischer Verbindungen. *Zeitschrift für Anorganische Chemie* **1893**, *3*, 267–330.

- (2) Kallies, B.; Meier, R. Electronic Structure of 3d $M(H_2O)_6^{3+}$ Ions from Sc^{III} to Fe^{III} : A Quantum Mechanical Study Based on DFT Computations and Natural Bond Orbital Analyses. *Inorg. Chem.* **2001**, *40*, 3101–3112.
- (3) Erras-Hanauer, H.; Clark, T.; van Eldik, R. Molecular Orbital and DFT Studies on Water Exchange Mechanisms of Metal Ions. *Coord. Chem. Rev.* **2003**, *238*, 233–253.
- (4) Helm, L.; Merbach, A. E. Water Exchange on Metal Ions: Experiments and Simulations. *Coord. Chem. Rev.* **1999**, *187*, 151–181.
- (5) Iuchi, S.; Morita, A.; Kato, S. Potential Energy Surfaces and Dynamics of Ni^{2+} Ion Aqueous Solution: Molecular Dynamics Simulation of the Electronic Absorption Spectrum. *J. Chem. Phys.* **2004**, *121*, 8446–8457.
- (6) Deguillaume, L.; Leriche, M.; Desboeufs, K.; Mailhot, G.; George, C.; Chaumerliac, N. Transition Metals in Atmospheric Liquid Phases: Sources, Reactivity, and Sensitive Parameters. *Chem. Rev.* **2005**, *105*, 3388–3431.
- (7) Ohtaki, H.; Radnai, T. Structure and Dynamics of Hydrated Ions. *Chem. Rev.* **1993**, *93*, 1157–1204.
- (8) Figgis, B. N.; Hitchman, M. A. *Ligand Field Theory and Its Applications*; Wiley-VCH, 2000.
- (9) Griffith, J. S. *The Theory of Transition-Metal Ions*; Cambridge University Press, 1961.
- (10) Tanabe, Y.; Sugano, S. On the Absorption Spectra of Complex Ions .1. *J. Phys. Soc. Jap.* **1954**, *9*, 753–766.
- (11) Tanabe, Y.; Sugano, S. On the Absorption Spectra of Complex Ions .2. *J. Phys. Soc. Jap.* **1954**, *9*, 766–779.
- (12) Stöhr, J. *NEXAFS Spectroscopy*; Springer: Berlin, 1992.

- (13) Aziz, E. F. The Solvation of Ions and Molecules Probed via Soft X-ray Spectroscopies. *J. Electron Spectrosc. Relat. Phenom.* **2010**, *177*, 168–180.
- (14) Aziz, E. F.; Eisebitt, S.; de Groot, F.; Chiou, J. W.; Dong, C. G.; Guo, J. H.; Eberhardt, W. Direct Contact versus Solvent-Shared Ion Pairs in NiCl₂ Electrolytes Monitored by Multiplet Effects at Ni(II) L Edge X-ray Absorption. *J. Phys. Chem. B* **2007**, *111*, 4440–4445.
- (15) Aziz, E. F.; Rittmann-Frank, M. H.; Lange, K. M.; Bonhommeau, S.; Chergui, M. Charge Transfer to Solvent Identified Using Dark Channel Fluorescence-Yield L-edge Spectroscopy. *Nat. Chem.* **2010**, *2*, 853–857.
- (16) Bonhommeau, S.; Ottosson, N.; Pokapanich, W.; Svensson, S.; Eberhardt, W.; Björneholm, O.; Aziz, E. F. Solvent Effect of Alcohols at the L-edge of Iron in Solution: X-ray Absorption and Multiplet Calculations. *J. Phys. Chem. B* **2008**, *112*, 12571–12574.
- (17) Schreck, S.; Gavrilu, G.; Weniger, C.; Wernet, Ph. A Sample Holder for Soft X-ray Absorption Spectroscopy of Liquids in Transmission Mode. *Rev. Sci. Instrum.* **2011**, *82*, 103101.
- (18) Näslund, L. A.; Cavalleri, M.; Ogasawara, H.; Nilsson, A.; Pettersson, L. G. M.; Wernet, Ph.; Edwards, D. C.; Sandström, M.; Myneni, S. Direct Evidence of Orbital Mixing Between Water and Solvated Transition-Metal Ions: An Oxygen 1s XAS and DFT Study of Aqueous Systems. *J. Phys. Chem. A* **2003**, *107*, 6869–6876.
- (19) Gotz, M. D.; Soldatov, M. A.; Lange, K. M.; Engel, N.; Golnak, R.; Könnecke, R.; Atak, K.; Eberhardt, W.; Aziz, E. F. Probing Coster-Kronig Transitions in Aqueous Fe²⁺ Solution Using Inverse Partial and Partial Fluorescence Yield at the L-Edge. *J. Phys. Chem. Lett.* **2012**, *3*, 1619–1623.
- (20) Huse, N.; Cho, H.; Hong, K.; Jamula, L.; de Groot, F. M. F.; Kim, T. K.; McCusker, J. K.; Schoenlein, R. W. Femtosecond Soft X-ray Spectroscopy of Solvated

- Transition-Metal Complexes: Deciphering the Interplay of Electronic and Structural Dynamics. *J. Phys. Chem. Lett.* **2011**, *2*, 880–884.
- (21) Huse, N.; Kim, T. K.; Jamula, L.; McCusker, J. K.; de Groot, F. M. F.; Schoenlein, R. W. Photo-Induced Spin-State Conversion in Solvated Transition Metal Complexes Probed via Time-Resolved Soft X-ray Spectroscopy. *J. Am. Chem. Soc.* **2010**, *132*, 6809–6816.
- (22) Soldatov, M. A.; Lange, K. M.; Gotz, M. D.; Engel, N.; Golnak, R.; Kothe, A.; Aziz, E. F. On the Origin of Dips in Total Fluorescence Yield X-ray Absorption Spectra: Partial and Inverse Partial Fluorescence Yield at the L-edge of Cobalt Aqueous Solution. *Chem. Phys. Lett.* **2012**, *546*, 164–167.
- (23) Winter, B.; Faubel, M. Photoemission from Liquid Aqueous Solutions. *Chem. Rev.* **2006**, *106*, 1176–1211.
- (24) Seidel, R.; Ghadimi, S.; Lange, K. M.; Bonhommeau, S.; Soldatov, M. A.; Golnak, R.; Kothe, A.; Könnecke, R.; Soldatov, A.; Thürmer, S. et al. Origin of Dark-Channel X-ray Fluorescence from Transition-Metal Ions in Water. *J. Am. Chem. Soc.* **2012**, *134*, 1600–1605.
- (25) Seidel, R.; Thürmer, S.; Moens, J.; Geerlings, P.; Blumberger, J.; Winter, B. Valence Photoemission Spectra of Aqueous $\text{Fe}^{2+/3+}$ and $\text{Fe}(\text{CN})_6^{4-/3-}$ and Their Interpretation by DFT Calculations. *J. Phys. Chem. B* **2011**, *115*, 11671–11677.
- (26) Seidel, R.; Thürmer, S.; Winter, B. Photoelectron Spectroscopy Meets Aqueous Solution: Studies from a Vacuum Liquid Microjet. *J. Phys. Chem. Lett.* **2011**, *2*, 633–641.
- (27) Thürmer, S.; Seidel, R.; Eberhardt, W.; Bradforth, S. E.; Winter, B. Ultrafast Hybridization Screening in Fe^{3+} Aqueous Solution. *J. Am. Chem. Soc.* **2011**, *133*, 12528–12535.

- (28) Lange, K. M.; Könnecke, R.; Ghadimi, S.; Golnak, R.; Soldatov, M. A.; Hodeck, K. F.; Soldatov, A.; Aziz, E. F. High Resolution X-ray Emission Spectroscopy of Water and Aqueous Ions Using the Micro-Jet Technique. *Chem. Phys.* **2010**, *377*, 1–5.
- (29) Lange, K. M.; Kothe, A.; Aziz, E. F. Chemistry in Solution: Recent Techniques and Applications Using Soft X-ray Spectroscopy. *Phys. Chem. Chem. Phys.* **2012**, *14*, 5331–5338.
- (30) Wernet, Ph.; Kunnus, K.; Schreck, S.; Quevedo, W.; Kurian, R.; Techert, S.; de Groot, F. M. F.; Odellius, M.; Föhlisch, A. Dissecting Local Atomic and Intermolecular Interactions of Transition-Metal Ions in Solution with Selective X-ray Spectroscopy. *J. Phys. Chem. Lett.* **2012**, *3*, 3448–3453.
- (31) Kunnus, K.; Rajkovic, I.; Schreck, S.; Quevedo, W.; Eckert, S.; Beye, M.; Suljoti, E.; Weniger, C.; Kalus, C.; Grübel, S. et al. A Setup for Resonant Inelastic Soft X-ray Scattering on Liquids at Free Electron Laser Light Sources. *Rev. Sci. Instrum.* **2012**, *83*, 123109.
- (32) Suljoti, E.; Garcia-Diez, R.; Bokarev, S. I.; Lange, K. M.; Schoch, R.; Dierker, B.; Dantz, M.; Yamamoto, K.; Engel, N.; Atak, K. et al. Direct Observation of Molecular Orbital Mixing in a Solvated Organometallic Complex. *Angew. Chem. Int. Ed.* **2013**, *52*, 98419844.
- (33) Atak, K.; Bokarev, S. I.; Gotz, M.; Golnak, R.; Lange, K. M.; Engel, N.; Dantz, M.; Suljoti, E.; Kühn, O.; Aziz, E. F. Nature of the Chemical Bond of Aqueous Fe²⁺ Probed by Soft X-ray Spectroscopies and ab Initio Calculations. *J. Phys. Chem. B* **2013**, DOI: 10.1021/jp408212u.
- (34) Ament, L. J. P.; van Veenendaal, M.; Devereaux, T. P.; Hill, J. P.; van den Brink, J. Resonant Inelastic X-ray Scattering Studies of Elementary Excitations. *Rev. Mod. Phys.* **2011**, *83*, 705–767.

- (35) de Groot, F. M. F. High Resolution X-ray Emission and X-ray Absorption Spectroscopy. *Chem. Rev.* **2001**, *101*, 1779–1808.
- (36) Föhlisch, A.; Hasselström, J.; Bennich, P.; Wassdahl, N.; Karis, O.; Nilsson, A.; Triguero, L.; Nyberg, M.; Pettersson, L. G. M. Ground-State Interpretation of X-ray Emission Spectroscopy on Adsorbates: CO Adsorbed on Cu(100). *Phys. Rev. B* **2000**, *61*, 16229–16240.
- (37) Ågren, H.; Luo, Y.; Gel'mukhanov, F.; Jensen, H. J. A. Screening in Resonant X-ray Emission of Molecules. *J. Electron Spectrosc. Relat. Phenom.* **1996**, *82*, 125–134.
- (38) Gel'mukhanov, F.; Ågren, H. Raman, Non-Raman, and Anti-Raman Dispersion in Resonant X-ray Scattering Spectra of Molecules. *Phys. Rev. A* **1996**, *54*, 3960–3970.
- (39) Gel'mukhanov, F.; Ågren, H. Resonant X-ray Raman Scattering. *Phys. Rep.* **1999**, *312*, 87–330.
- (40) de Groot, F. M. F.; Arrio, M. A.; Saintavit, P.; Cartier, C.; Chen, C. T. Fluorescence Yield Detection: Why It Does Not Measure the X-ray Absorption Cross Section. *Solid State Comm.* **1994**, *92*, 991–995.
- (41) Ghiringhelli, G.; Matsubara, M.; Dallera, C.; Fracassi, F.; Gusmeroli, R.; Piazzalunga, A.; Tagliaferri, A.; Brookes, N. B.; Kotani, A.; Braicovich, L. NiO As a Test Case for High Resolution Resonant Inelastic Soft X-ray Scattering. *J. Phys.: Condens. Matter* **2005**, *17*, 5397–5412.
- (42) de Groot, F. M. F.; Fuggle, J. C.; Thole, B. T.; Sawatzky, G. A. 2p X-Ray Absorption of 3d Transition-Metal Compounds - An Atomic Multiplet Description Including the Crystal-Field. *Phys. Rev. B* **1990**, *42*, 5459–5468.
- (43) Thole, B. T.; van der Laan, G.; Fuggle, J. C.; Sawatzky, G. A.; Karnatak, R. C.; Es-

- teva, J.-M. 3d X-ray Absorption Lines and the $3d^94f^{m+1}$ Multiplets of the Lanthanides. *Phys. Rev. B* **1985**, *32*, 5107–5118.
- (44) Haverkort, M. W.; Zwierzycki, M.; Andersen, O. K. Multiplet Ligand-Field Theory Using Wannier Orbitals. *Phys. Rev. B* **2012**, *85*, 165113.
- (45) Agui, A.; Uozumi, T.; Mizumaki, M.; Käämbre, T. Intermetallic Charge Transfer in FeTiO₃ Probed by Resonant Inelastic Soft X-ray Scattering. *Phys. Rev. B* **2009**, *79*, 092402.
- (46) Roemelt, M.; Maganas, D.; DeBeer, S.; Neese, F. A Combined DFT and Restricted Open-Shell Configuration Interaction Method Including Spin-Orbit Coupling: Application to Transition Metal L-edge X-ray Absorption Spectroscopy. *J. Chem. Phys.* **2013**, *138*, 204101.
- (47) Roemelt, M.; Neese, F. Excited States of Large Open-Shell Molecules: An Efficient, General, and Spin-Adapted Approach Based on a Restricted Open-Shell Ground State Wave function. *J. Phys. Chem. A* **2013**, *117*, 3069–3083.
- (48) Ikeno, H.; Tanaka, I.; Koyama, Y.; Mizoguchi, T.; Ogasawara, K. First-Principles Multielectron Calculations of Ni L₂,L₃ NEXAFS and ELNES for LiNiO₂ and Related Compounds. *Phys. Rev. B* **2005**, *72*, 075123.
- (49) Ikeno, H.; Mizoguchi, T.; Tanaka, I. *Ab initio* Charge Transfer Multiplet Calculations on the L₂,L₃ XANES and ELNES of 3d Transition Metal Oxides. *Phys. Rev. B* **2011**, *83*, 155107.
- (50) Vinson, J.; Rehr, J. J.; Kas, J. J.; Shirley, E. L. Bethe-Salpeter Equation Calculations of Core Excitation Spectra. *Phys. Rev. B* **2011**, *83*, 115106.
- (51) Vinson, J.; Rehr, J. J. *Ab initio* Bethe-Salpeter Calculations of the X-ray Absorption Spectra of Transition Metals at the L-Shell Edges. *Phys. Rev. B* **2012**, *86*, 195135.

- (52) Josefsson, I.; Kunnus, K.; Schreck, S.; Föhlisch, A.; de Groot, F. M. F.; Wernet, Ph.; Odellius, M. *Ab initio* Calculations of X-ray Spectra: Atomic Multiplet and Molecular Orbital Effects in a Multiconfigurational SCF Approach to the L-Edge Spectra of Transition Metal Complexes. *J. Phys. Chem. Lett.* **2012**, *3*, 3565–3570.
- (53) Bokarev, S. I.; Dantz, M.; Suljoti, E.; Kühn, O.; Aziz, E. F. State-Dependent Electron Delocalization Dynamics at the Solute-Solvent Interface: Soft-X-Ray Absorption Spectroscopy and AbInitio Calculations. *Phys. Rev. Lett.* **2013**, *111*, 083002.
- (54) Magini, M. Hydration and Complex Formation Study on Concentrated MCl₂ Solutions [M=Co(II), Ni(II), Cu(II)] by X-ray Diffraction Technique. *J. Chem. Phys.* **1981**, *74*, 2523–2529.
- (55) Waizumi, K.; Kouda, T.; Tanio, A.; Fukushima, N.; Ohtaki, H. Structural Studies on Saturated Aqueous Solutions of Manganese(II), Cobalt(II), and Nickel(II) Chlorides by X-ray Diffraction. *J. Solution Chem.* **1999**, *28*, 83–100.
- (56) Aquilante, F.; De Vico, L.; Ferre, N.; Ghigo, G.; Malmqvist, P. A.; Neogrady, P.; Pedersen, T. B.; Pitonak, M.; Reiher, M.; Roos, B. O. et al. Software News and Update MOLCAS 7: The Next Generation. *J. Comput. Chem.* **2010**, *31*, 224–247.
- (57) Roos, B. O.; Lindh, R.; Malmqvist, P. A.; Veryazov, V.; Widmark, P. O. Main Group Atoms and Dimers Studied with a New Relativistic ANO Basis Set. *J. Phys. Chem. A* **2004**, *108*, 2851–2858.
- (58) Roos, B. O.; Lindh, R.; Malmqvist, P. A.; Veryazov, V.; Widmark, P. O. New Relativistic ANO Basis Sets for Transition Metal Atoms. *J. Phys. Chem. A* **2005**, *109*, 6575–6579.
- (59) Douglas, M.; Kroll, N. M. Quantum Electrodynamical Corrections to Fine-Structure of Helium. *Annals of Physics* **1974**, *82*, 89–155.

- (60) Hess, B. A. Relativistic Electronic-Structure Calculations Employing a 2-component No-Pair Formalism with External-Field Projection Operators. *Phys. Rev. A* **1986**, *33*, 3742–3748.
- (61) Malmqvist, P. A.; Rendell, A.; Roos, B. O. The Restricted Active Space Self-Consistent-Field Method, Implemented with a Split Graph Unitary-Group Approach. *J. Phys. Chem.* **1990**, *94*, 5477–5482.
- (62) Finley, J.; Malmqvist, P. A.; Roos, B. O.; Serrano-Andres, L. The Multi-State CASPT2 Method. *Chem. Phys. Lett.* **1998**, *288*, 299–306.
- (63) Malmqvist, P. A.; Roos, B. O.; Schimmelpfennig, B. The Restricted Active Space (RAS) State Interaction Approach with Spin-Orbit Coupling. *Chem. Phys. Lett.* **2002**, *357*, 230–240.
- (64) Kurian, R.; Kunnus, K.; Wernet, Ph.; Butorin, S. M.; Glatzel, P.; de Groot, F. M. F. Intrinsic Deviations in Fluorescence Yield Detected X-ray Absorption Spectroscopy: the Case of the Transition Metal $L_{2,3}$ Edges. *J. Phys.: Condens. Matter* **2012**, *24*, 452201.
- (65) Krause, M. Atomic Radiative and Radiationless Yield for K and L Shells. *J. Phys. Chem. Ref. Data* **1979**, *8*, 307–327.
- (66) Kunnus, K.; Josefsson, I.; Schreck, S.; Quevedo, W.; Miedema, P. S.; Techert, S.; de Groot, F. M. F.; Föhlich, A.; Odellius, M.; Wernet, Ph. unpublished.
- (67) Coster, D.; Kronig, R. D. L. New type of auger effect and its influence on the x-ray spectrum. *Physica* **1935**, *2*, 13 – 24.
- (68) Casanova, D.; Head-Gordon, M. Restricted Active Space Spin-Flip Configuration Interaction Approach: Theory, Implementation and Examples. *Phys. Chem. Chem. Phys.* **2009**, *11*, 9779–9790.

- (69) Landry-Hum, J.; Bussiere, G.; Daniel, C.; Reber, C. Triplet Electronic States in d^2 and d^8 Complexes Probed by Absorption Spectroscopy: A CASSCF/CASPT2 Analysis of $V(H_2O)_6^{3+}$ and $Ni(H_2O)_6^{2+}$. *Inorg. Chem.* **2001**, *40*, 2595–2601.
- (70) de Groot, F.; Kotani, A. *Spectroscopy of Solids*; CRC Press, 2008.
- (71) de Groot, F. M. F. Multiplet Effects in X-ray Spectroscopy. *Coord. Chem. Rev.* **2005**, *249*, 31–63.
- (72) van Schooneveld, M. M.; Gosselink, R. W.; Eggenhuisen, T. M.; Al Samarai, M.; Monney, C.; Zhou, K. J.; Schmitt, T.; de Groot, F. M. F. A Multispectroscopic Study of 3d Orbitals in Cobalt Carboxylates: The High Sensitivity of 2p3d Resonant X-ray Emission Spectroscopy to the Ligand Field. *Angew. Chem. Int. Ed.* **2013**, *52*, 1170–1174.